

## Whitewash and the limestone cycle

Limestone (*chaux*) containing calcium carbonate,  $\text{CaCO}_3$ , burned to remove carbon dioxide, yields Quicklime (*chaux vive*, calcium oxide,  $\text{CaO}$ ). This reacts with water to produce hydrated or 'slaked' lime (*chaux éteinte*, calcium hydroxide,  $\text{Ca(OH)}_2$ ). Limewater is a solution of this in water and can be used as whitewash. Exposure to carbon dioxide will return this to limestone.

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